**Gas Laws Quiz Review – Honors**

1. Gases behave most ideally at low pressures and least ideally at high pressures. Given what you know of the kinetic molecular theory, explain why this might be.
2. Explain why it is reasonable to say that the molecules of an ideal gas do not experience intermolecular forces.
3. If I have 15 grams of methane gas in a container with a volume of 2.75 liters at a temperature of 25 degrees Celsius, what will the pressure inside the container be?
4. If I were to connect the container from problem 3 to an empty container with a volume of 1.75 L, what would the new pressure of the methane be after it rushed into the other container?
5. Define “pressure.”
6. Why does heating a gas cause the pressure of the container it is in to increase?